

Heating Curve Physics

latent heat and specific heat capacity questions. - $c_{\text{water}} = 4200 \text{ J/kgK}$ $c_{\text{ice}} = 2100 \text{ J/kgK}$ $c_{\text{copper}} = 390 \text{ J/kgK}$ L_f , L_v , $L_s = 3.34 \times 10^5 \text{ J/kg}$ $L_v = 22.5 \times 10^5 \text{ J/kg}$ latent heat and specific heat capacity questions. 1. how much water at 50°C is needed to just melt 2.2 kg of ice at 0°C ?

earth energy designer - building physics - software within building physics and ground heat storage blocon buildingphysics eed v3.2 earth energy designer february 19, 2015 **solubility of a salt - chymist** - table 1. quantities of salt and water add another portion of water (see table 1 for the amount). to do this, measure the next quantity of water, as listed in table 1, in a 10 ml graduated cylinder (assume that 1 g of water has a volume of 1 ml)

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